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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{\rho}$  32 g mol / 1.790 g l = 17.86 l

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as  $6.02 \times 10^{23}$  particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions

a mole is always  $6.02 \times 10^{23}$  but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole m<sup>3</sup> mol although it is more typical to use the units cubic decimetres per mole dm<sup>3</sup> mol for gases and cubic centimetres per mole cm<sup>3</sup> mol for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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temperature pressure and volume are discussed the  
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molecular level and simple models are introduced that allow  
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 $V = 22.4 \text{ l}$  in other words 1 mole of a gas will occupy 22 4 l at  
stp assuming ideal gas behavior at stp the volume of a gas is  
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